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chapter problems.

Due Dates. 1.

Describe properties of
acids and bases 2.

Compare and contrast
the theories of
Arrhenius and Br ø

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- 49, 72, 74, 75: 5/29:

3. Describe how H^+
and OH^-

concentrations are
related in an aqueous
solution. 4.

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Describe properties of
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3. Describe how H^+
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solution. 4. 19

Acids Bases

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Section 19.3

Strengths of Acids

and Bases □

OBJECTIVES:

□ Calculate an acid dissociation constant (K_a) from concentration and pH measurements.

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Chapter 19.3

Strengths of Acids and Bases □

OBJECTIVES: □ Order acids by strength according to their acid dissociation constants (K_a).

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Bases and Salts Acid-
Base Theories ...

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Chapter 19
Acids, Bases, and
Salts Chemistry
Section 19.1 Acid-
Base Theories

OBJECTIVES:

Define the properties
of acids and bases.

Section 19.1 Acid-
Base Theories

OBJECTIVES:

Compare and
contrast acids and
bases as defined by
the theories of: a)

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Arrhenius, b)
Brønsted-Lowry, and
c) Lewis.

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□ Acids, Bases, and
Salts □ Chemistry
Section 19.1 Acid-
Base Theories
OBJECTIVES:

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Chapter 19
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Objectives
Answers

□ Define the properties of acids and bases.

Section 19.1 Acid-Base Theories

OBJECTIVES:

□ Compare and contrast acids and bases as defined by the theories of: a) Arrhenius, b) Brønsted-Lowry, and c) Lewis.

Chemistry - Chapter

Page 20/39

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Salts - Chapter ...

In writing nucleotide sequences for nucleic acids, the convention is to write the nucleotides (usually using the one-letter abbreviations for the bases, shown in Figure 19.5 □ Structure of a Segment of DNA □) starting with the nucleotide having

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a free phosphate group, which is known as the 5' end, and indicate the nucleotides in order.

Answers

19.2 Nucleic Acid
Structure | The Basics
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OBJECTIVES: Order acids by strength according to their acid dissociation constants (K_a). Section 19.3 Strengths of Acids and Bases

OBJECTIVES: Order bases by strength according to their base dissociation constants (K_b). Strength. Acids and Bases are classified

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According to the degree to which they ionize in water: Strong acids are completely ionized in aqueous solution; this means they ionize 100 %. Weak acids ionize only slightly in aqueous solution. Strength

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19: Bases and Acids.

STUDY. PLAY. Acids.

taste sour, change the
color of an acid base
indicator, can be

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strong or weak electrolytes in aqueous solution Ex: citrus. bases. taste bitter, feel slippery, will change the color of an acid base indicator, can be strong or weak electrolytes in aqueous solution. Ex: soap.
Arrhenius Acids.

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distribution of Earth's

water. Identify water's

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structure and
properties. Define
acids, bases, and pH.
Objectives
Explain why water is
essential for life.

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pages. Vocabulary.
acid solution with a
pH lower than 7; base
solution with a pH
higher than 7;
hydrogen bond

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OBJECTIVES

: Define the properties
of acids and bases.

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Section 19.1 Acid-
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OBJECTIVES

- Compare and contrast acids and bases as defined by the theories of: a) Arrhenius, b) Brønsted-Lowry, and c) Lewis.

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